

$$V(\text{CO}_2) = V_1(\text{CO}_2) + V_2(\text{CO}_2) = 1,4x + 1,493 \cdot (28 - x) = 41,44$$

$$1,4x + 41,804 - 1,493x = 41,44$$

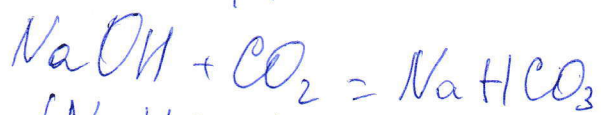
$$x = 4$$

$$\text{CH}_4 = 4:28 \cdot 100\% = 14,3\%$$

$$\text{C}_2\text{H}_6 = 100\% - 14,3\% = 85,7\% \quad 65$$

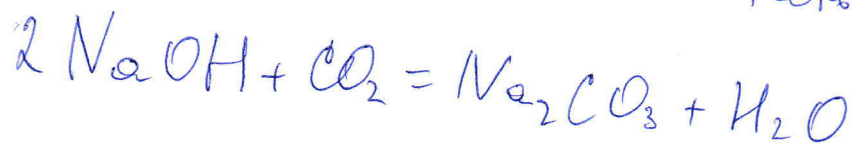
Вопрос 3

$$n(\text{CO}_2) = \frac{10}{22,4} = 0,446 \text{ моль}$$



$$n(\text{NaHCO}_3) = 0,446 \text{ моль}$$

$$m(\text{NaHCO}_3) = 0,446 \text{ моль} \cdot 84 \frac{\text{г}}{\text{моль}} = 37,5$$



$$n(\text{Na}_2\text{CO}_3) = 0,375 \text{ моль}$$

$$m(\text{Na}_2\text{CO}_3) = 0,375 \cdot 106 = 39,75 \text{ г}$$

$$m(\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}) = 0,375 \text{ моль} \cdot 286 \frac{\text{г}}{\text{моль}} = 107,25 \text{ г}$$